

## TITLE OF THE INVENTION

NON-AQUEOUS ELECTROLYTE BATTERY

## BACKGROUND OF THE INVENTION

### Field of the Invention

[0001]

The present invention relates generally to a non-aqueous electrolyte battery provided with a positive electrode, a negative electrode, and a non-aqueous electrolyte using an organic solvent, and particularly, is characterized in that a non-aqueous electrolyte battery having a high battery capacity is obtained upon improvement of a positive electrode material used for the positive electrode.

### Description of the Related Art

[0002]

Recently, as one type of advanced batteries featuring high power and high energy density, a non-aqueous electrolyte battery has come into practical use, wherein a non-aqueous electrolyte solution using an organic solvent is employed and discharge/charge is performed by way of transferring lithium ions between a positive electrode and a negative electrode.

[0003]

In such a non-aqueous electrolyte battery, a lithium-transition metal compound oxide capable of occluding and discharging lithium such as  $\text{LiCoO}_2$  and the like has been

generally used as a positive electrode material for the positive electrode, and a carbon material such as graphite and the like as a negative electrode material for the negative electrode.

[0 0 0 4]

In a non-aqueous electrolyte battery as described above, a specific capacity of graphite as the negative electrode material is about 370 Ah/kg. On the other hand, when  $\text{LiCoO}_2$  is used as the positive electrode material, Li contained in  $\text{LiCoO}_2$  is not sufficiently discharged, and accordingly, a specific capacity of  $\text{LiCoO}_2$  is as low as 150 Ah/kg. Energy density of such a non-aqueous electrolyte battery is about 150 Wh/kg.

[0 0 0 5]

Recently, however, such a non-aqueous electrolyte battery has come into practical use as an electric current source of a cordless telephone, a personal computer, and the like, and accordingly, it has been desired that a high battery capacity of the above-mentioned non-aqueous electrolyte battery is obtained upon further improvement of energy density thereof.

[0 0 0 6]

Further, in recent years, it has been discussed that  $\text{Fe}_2\text{O}_3$  and  $\text{FeS}_2$  are used as the positive electrode material of the non-aqueous electrolyte battery. However,  $\text{Fe}_2\text{O}_3$  and  $\text{FeS}_2$  have

been used as a positive electrode material and the like only for a primary battery and a high-temperature battery, that is, have not been utilized effectively in a non-aqueous electrolyte battery provided with a non-aqueous electrolyte using an organic solvent.

#### SUMMARY OF THE INVENTION

[0007]

An object of the present invention is, in a non-aqueous electrolyte battery provided with a positive electrode, a negative electrode, and a non-aqueous electrolyte using an organic solvent, to improve a specific capacity of a positive electrode material used for a positive electrode, thereby obtaining a non-aqueous electrolyte battery having a high battery capacity.

[0008]

A non-aqueous electrolyte battery according to a first aspect of the present invention is provided with a positive electrode, a negative electrode, and a non-aqueous electrolyte using an organic solvent, and uses for the positive electrode a positive electrode material containing at least ferrite and as a negative electrode material for the negative electrode a lithium alloy.

[0009]

A non-aqueous electrolyte battery according to a second

aspect of the present invention is provided with a positive electrode, a negative electrode, and a non-aqueous electrolyte using an organic solvent, and uses for the positive electrode a positive electrode material containing at least  $\text{FeS}_2$ , and as a negative electrode material for the negative electrode a lithium alloy.

[0010]

In the non-aqueous electrolyte batteries according to the first and second aspects of the present invention, discharge is performed first and foremost so as to intercalate lithium contained in the negative electrode material composed of a lithium alloy used for the negative electrode into the positive electrode material composed of ferrite or  $\text{FeS}_2$  used for the positive electrode, after which, lithium ions are transferred between the positive electrode and the negative electrode, to perform discharge/charge.

[0011]

When discharge is performed first and foremost so as to intercalate lithium contained in a lithium alloy as the negative electrode material into ferrite or  $\text{FeS}_2$  as the positive electrode material as described above, a large amount of lithium is contributed to discharge/charge, thereby improving a specific capacity of the positive electrode material, resulting in increased energy density of the non-aqueous electrolyte battery. Therefore, a non-aqueous electrolyte

battery having a high battery capacity is obtained.

[0 0 1 2]

In the above-mentioned non-aqueous electrolyte battery according to the first aspect of the present invention, examples of usable ferrite as the positive electrode material include  $\text{Fe}_2\text{O}_3$ ,  $\text{Fe}_3\text{O}_4$ ,  $\text{MnFe}_2\text{O}_4$ ,  $\text{NiFe}_2\text{O}_4$ ,  $\text{CoFe}_2\text{O}_4$ ,  $\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$ , and the like. In particular, in the case where  $\text{CoFe}_2\text{O}_4$  is used as the positive electrode material, lithium contained in a lithium alloy as the negative electrode material is intercalated into  $\text{CoFe}_2\text{O}_4$  by way of the initial discharge, after which, lithium is properly discharged from  $\text{CoFe}_2\text{O}_4$  during charging, thereby improving extremely a specific capacity of the positive electrode material, resulting in remarkably increased energy density. Therefore, the non-aqueous electrolyte battery having a high battery capacity is obtained.

[0 0 1 3]

In the non-aqueous electrolyte batteries according to the first and second aspects of the present invention, examples of usable lithium alloy as the negative electrode material include lithium alloy for all types. More specifically, it is preferable that a lithium alloy (Li-Si) having a high specific capacity is used.

[0 0 1 4]

Further, a non-aqueous electrolyte battery according to a third aspect of the present invention is provided with a

positive electrode, a negative electrode, and a non-aqueous electrolyte using an organic solvent, and uses as a positive electrode material for the positive electrode a transition metal oxide (except for  $\text{LiCoO}_2$ ) having crystal structure of space group R3m and for the negative electrode a negative electrode material containing lithium.

[0015]

In the non-aqueous electrolyte battery according to the third aspect of the present invention, discharge is performed first and foremost so as to intercalate lithium contained in the negative electrode material for the negative electrode into the positive electrode material using the transition metal oxide (except for  $\text{LiCoO}_2$ ) having the crystal structure of space group R3m, after which, lithium ions are transferred between the positive electrode and the negative electrode, to perform discharge/charge, as shown above in the non-aqueous electrolyte batteries according to the first and second aspects of the present invention.

[0016]

When discharge is performed first and foremost so as to intercalate lithium contained in the negative electrode material for the negative electrode into the positive electrode material using the transition metal oxide (except for  $\text{LiCoO}_2$ ) having the crystal structure of space group R3m, as described above, a large amount of lithium is contributed to

discharge/charge, thereby improving a specific capacity of the positive electrode material, resulting in increased energy density of the non-aqueous electrolyte battery. Therefore, the non-aqueous electrolyte battery having a high battery capacity is obtained.

[0 0 1 7]

Examples of usable transition metal oxide having the crystal structure of space group R3m include  $\text{NaFeO}_2$ ,  $\text{LiNiO}_2$ ,  $\text{LiCoO}_2$ , and the like. However, in the case where  $\text{LiCoO}_2$  is used as the positive electrode material, when lithium contained in the negative electrode material is intercalated into  $\text{LiCoO}_2$  by way of the aforementioned discharge, lithium is not discharged from  $\text{LiCoO}_2$  during the charging after the intercalation. It is supposed that the reason why lithium is not discharged from  $\text{LiCoO}_2$  as shown above, is that the intercalation of lithium changes the crystal structure of  $\text{LiCoO}_2$ .

[0 0 1 8]

On the other hand, in the case where transition metal oxides other than  $\text{LiCoO}_2$  having the crystal structure of space group R3m are used, when lithium contained in the negative electrode material is intercalated into the positive electrode material by way of the aforementioned discharge, and then charge is performed, lithium is properly discharged from the positive electrode material, resulting in the improved

specific capacity of the positive electrode material. In particular, in the case where  $\text{NaFeO}_2$  is used for the positive electrode material, the specific capacity of the positive electrode material is extremely improved, resulting in remarkably increased energy density. Therefore, the non-aqueous electrolyte battery having a high battery capacity is obtained.

[0019]

Furthermore, in the non-aqueous electrolyte battery according to the third aspect of the present invention, any material can be used as the negative electrode material for the negative electrode as long as the material can contain lithium and serve to supply lithium into the above-mentioned positive electrode material by way of the aforementioned discharge as describe above. Examples of usable material include lithium metal and lithium alloy for all types. In addition, a carbon material and the like which have lithium occluded therein can be used. More specifically, it is preferable that a lithium alloy (Li-Si) having a high specific capacity can be used.

[0020]

Moreover, in the non-aqueous electrolyte batteries according to the first to third aspects of the present invention, as the non-aqueous electrolyte using the organic solvent, ones that have been generally used can be employed.



[0021]

As the organic solvent used for the non-aqueous electrolyte, well-known ones that have been generally used in a non-aqueous electrolyte solution can be employed. More specifically, cyclic carbonic ester such as ethylene carbonate, propylene carbonate, butylene carbonate, vinylene carbonate, and the like, and chain carbonic ester such as dimethyl carbonate, diethyl carbonate, dipropyl carbonate, ethyl methyl carbonate, ethyl propyl carbonate, and the like, can be used alone or in combination of two or more types.

[0022]

Further, in the non-aqueous electrolyte, as a solute being dissolved in the above-mentioned organic solvent, well-known ones can be used. More specifically, a lithium compound such as trifluoro-methanesulfonic acid lithium ( $\text{LiCF}_3\text{SO}_3$ ), hexafluorophosphoric acid lithium ( $\text{LiPF}_6$ ), perchloric acid lithium ( $\text{LiClO}_4$ ), tetrafluoroboric acid lithium ( $\text{LiBF}_4$ ), trifluoro-methanesulfonic acid imide lithium  $\{\text{LiN}(\text{CF}_3\text{SO}_2)_2\}$ , and the like can be used.

[0023]

Moreover, in the non-aqueous electrolyte batteries according to the first to third aspects of the present invention, as a separator separating the positive electrode and the negative electrode, well-known ones which have been generally used can be used.

[0 0 2 4]

These and other objects, advantages and features of the invention will become apparent from the following description thereof taken in conjunction with the accompanying drawings which illustrate specific embodiment of the invention.

#### BRIEF DESCRIPTION OF THE DRAWINGS

[0 0 2 5]

Fig.1 is a schematic explanatory view illustrating a test cell fabricated in each of Examples according to the present invention;

[0 0 2 6]

Fig.2 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example A1 of the present invention wherein  $\text{Fe}_2\text{O}_3$  is used as a positive electrode material;

[0 0 2 7]

Fig.3 is a diagram illustrating the characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example A1;

[0 0 2 8]

Fig.4 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example A2 of the present invention wherein  $\text{Fe}_3\text{O}_4$  is used as a positive electrode material;

[0 0 2 9]

Fig.5 is a diagram illustrating characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example A2;

[0 0 3 0]

Fig.6 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example A3 of the present invention wherein  $\text{MnFe}_2\text{O}_4$  is used as a positive electrode material;

[0 0 3 1]

Fig.7 is a diagram illustrating characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example A3;

[0 0 3 2]

Fig.8 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example A4 of the present invention wherein  $\text{NiFe}_2\text{O}_4$  is used as a positive electrode material;

[0 0 3 3]

Fig.9 is a diagram illustrating characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example A4;

[0 0 3 4]

Fig.10 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to

Example A5 of the present invention wherein  $\text{CoFe}_2\text{O}_4$  is used as a positive electrode material;

[0 0 3 5]

Fig.11 is a diagram illustrating characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example A5;

[0 0 3 6]

Fig.12 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example A6 of the present invention wherein  $\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$  is used as a positive electrode material;

[0 0 3 7]

Fig.13 is a diagram illustrating characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example A6;

[0 0 3 8]

Fig.14 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example B1 of the present invention wherein  $\text{FeS}_2$  is used as a positive electrode material;

[0 0 3 9]

Fig.15 is a diagram illustrating characteristics of discharge/charge at a first cycle in a test cell according to Example C1 of the present invention wherein  $\text{NaFeO}_2$  is used as a positive electrode material;

[0 0 4 0]

Fig.16 is a diagram illustrating characteristics of discharge/charge at a second cycle in the test cell according to the above-mentioned Example C1;

[0 0 4 1]

Fig.17 is a diagram illustrating a relationship between the number of cycles and a specific capacity of a positive electrode material in the test cell according to the above-mentioned Example C1; and,

[0 0 4 2]

Fig.18 is a diagram illustrating a relationship between the number of cycles and discharge/charge efficiency in the test cell according to the above-mentioned Example C1.

#### DESCRIPTION OF THE PREFERRED EMBODIMENTS

[0 0 4 3]

Next, the non-aqueous electrolyte battery according to the present invention will be described in detail by way of reference to Examples of the present invention. Furthermore, the description will also make apparent by way of comparison that in the non-aqueous electrolyte batteries according to Examples, a specific capacity of a positive electrode material is improved, resulting in increased energy density, whereby obtaining a high battery capacity. It is to be distinctly appreciated that the non-aqueous electrolyte battery

according to the present invention should not be limited to the following Examples but appropriate changes and modifications may be made in carrying out the present invention without departing from the spirit and scope of the present invention.

(Examples A1 to A6)

[0 0 4 4]

In each of Examples A1 to A6, each of positive electrodes and a non-aqueous electrolyte solution that were prepared as follows were used.

(Preparation of Positive Electrode)

[0 0 4 5]

In the preparation of a positive electrode, in Example A1,  $\text{Fe}_2\text{O}_3$  was used as a positive electrode material. Further,  $\text{Fe}_2\text{O}_3$  as the positive electrode material, acetylene black as a conductive agent, and polytetrafluoroethylene as a binding agent were mixed with each other in a ratio of 40 to 40 to 20 by weight, were formed by pressure into a disk of a diameter of 16 mm and a thickness of 0.1 mm. The resultant disk was dried in vacuum at  $110^\circ\text{C}$ , thereby to give the positive electrode of Example A1.

[0 0 4 6]

Besides, in Examples A2 to A6, each of positive electrodes was prepared in the same manner as in the above-mentioned Example A1 except that the type of the positive

electrode material in Example A1 was changed. More specifically,  $\text{Fe}_3\text{O}_4$  is used as a positive electrode material in Example A2,  $\text{MnFe}_2\text{O}_4$  in Example A3,  $\text{NiFe}_2\text{O}_4$  in Example A4,  $\text{CoFe}_2\text{O}_4$  in Example A5, and  $\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$  in Example A6.

(Preparation of Non-aqueous Electrolyte Solution)

[0 0 4 7]

In the preparation of a non-aqueous electrolyte solution, ethylene carbonate and dimethyl carbonate were mixed in a volume ratio of 1 to 1, to give a mixture solvent, in which a solute of hexafluorophosphoric acid lithium ( $\text{LiPF}_6$ ) was dissolved in a concentration of 1 mol/kg.

[0 0 4 8]

Next, as shown in Figure 1, the non-aqueous electrolyte solution 14 prepared as described above was contained in a test cell 10 as well as each of the positive electrodes prepared as described above was used for a working electrode 11 and lithium metal was used for a counter electrode 12 as a negative electrode and for a reference electrode 13.

[0 0 4 9]

Further, each of test cells according to Examples A1 to A6 wherein each of the above-mentioned positive electrodes was used for a working electrode 11 was subject to discharge at a  $2 \text{ mA/cm}^2$  discharging current to a potential of the working electrode 11 to the reference electrode 13 of 0.5 V, followed by charge at a  $2 \text{ mA/cm}^2$  charging current to a potential of the

working electrode 11 to the reference electrode 13 of 4.0 V. Afterward, discharge/charge was repeatedly performed in a cycle of the above-mentioned discharge/charge process, so that 10 cycles of discharge/charge were performed.

[0 0 5 0]

With regard to the test cell according to Example A1 wherein  $\text{Fe}_2\text{O}_3$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 2. Additionally, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second charge is graphed out as a charge curved line described in a solid line in Figure 3.

[0 0 5 1]

Further, with regard to the test cell according to Example A2 wherein  $\text{Fe}_3\text{O}_4$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge



is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 4. Additionally, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second charge is graphed out as a charge curved line described in a solid line in Figure 5.

[0 0 5 2]

Moreover, with regard to the test cell according to Example A3 wherein  $\text{MnFe}_2\text{O}_4$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 6. Additionally, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive

electrode material at a second charge is graphed out as a charge curved line described in a solid line in Figure 7.

[0 0 5 3]

Continually, with regard to the test cell according to Example A4 wherein  $\text{NiFe}_2\text{O}_4$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 8. Additionally, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second charge is graphed out as a charge curved line described in a solid line in Figure 9.

[0 0 5 4]

Furthermore, with regard to the test cell according to Example A5 wherein  $\text{CoFe}_2\text{O}_4$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and specific capacity

(Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 10. Additionally, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second charge is graphed out as a charge curved line in a solid line in Figure 11.

[0 0 5 5]

In addition, with regard to the test cell according to Example A6 wherein  $K_{1.4}Fe_{11}O_{17}$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line in a solid line in Figure 12. Additionally, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a second charge is graphed out as a charge curved line described in a solid line in Figure 13.

[0 0 5 6]

Besides, in the above-mentioned case where discharge/charge was performed using each of the positive electrodes according to Examples A1 to A6, the specific capacity (Ah/kg) of each of the positive electrodes at the discharges of the first and tenth cycles as well as a ratio of charge capacity to discharge capacity at the tenth cycle, that is, discharge/charge efficiency (%) at the tenth cycle were found. The results are shown in the following Table 1.

[0 0 5 7]

(TABLE 1)

	positive electrode material	specific capacity (Ah/kg)		discharge/ charge efficiency at tenth cycle (%)
		first cycle	tenth cycle	
Example A1	$\text{Fe}_2\text{O}_3$	1531	125	94
Example A2	$\text{Fe}_3\text{O}_4$	1013	272	98
Example A3	$\text{MnFe}_2\text{O}_4$	777	154	100
Example A4	$\text{NiFe}_2\text{O}_4$	1370	384	98
Example A5	$\text{CoFe}_2\text{O}_4$	1379	478	99
Example A6	$\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$	1300	342	98

[ 0 0 5 8 ]

As apparent from the results, in the case where  $\text{Fe}_2\text{O}_3$ ,  $\text{Fe}_3\text{O}_4$ ,  $\text{MnFe}_2\text{O}_4$ ,  $\text{NiFe}_2\text{O}_4$ ,  $\text{CoFe}_2\text{O}_4$ , and  $\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$ , which are ferrite are used as the positive electrode material, the specific capacity of each of the positive electrode materials at the discharge of the first cycle is higher than that of about 150 Ah/kg of  $\text{LiCoO}_2$  which has been conventionally used as the positive electrode material. Further, it is also apparent that in Examples A2, A4, A5 and A6 wherein  $\text{Fe}_3\text{O}_4$ ,  $\text{NiFe}_2\text{O}_4$ ,  $\text{CoFe}_2\text{O}_4$ , and  $\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$  are used as the positive electrode material, the specific capacity of the positive electrode material at the discharge of the tenth cycle is higher than that of  $\text{LiCoO}_2$  which has been conventionally used. It is to be distinctly noted that, in Example 5 wherein  $\text{CoFe}_2\text{O}_4$  is used as the positive electrode material, the specific capacity of the positive electrode material is extremely higher.

[ 0 0 5 9 ]

Further, in Examples A1 to A6, discharge/charge efficiency at the tenth cycle was approximately 100 %, making it possible to charge and discharge stably.

[ 0 0 6 0 ]

Next, in each of the non-aqueous electrolyte batteries according to Examples A1 to A6 wherein  $\text{Fe}_2\text{O}_3$ ,  $\text{Fe}_3\text{O}_4$ ,  $\text{MnFe}_2\text{O}_4$ ,  $\text{NiFe}_2\text{O}_4$ ,  $\text{CoFe}_2\text{O}_4$ , and  $\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$  are used as the positive electrode material while a lithium-alloy (Li-Si) having the specific

capacity of 3000 Ah/kg which is disclosed in the Japanese Patent Laid-Open No. 2000-321200 and the Japanese Patent Laid-Open No. 2000-321201 is used as the negative electrode material, and in the non-aqueous electrolyte battery according to Comparative Example wherein  $\text{LiCoO}_2$  having the specific capacity of 150 Ah/kg is used as the positive electrode material while graphite having the specific capacity of 370 Ah/kg is used as the negative electrode material, capacity density (Ah/kg) at the discharge of the first cycle is found by way of the following formula (1). The lithium alloy may include Ag, Al, Au, Bi, Ca, Co, Ga, Ge, Hg, In, Ir, Mg, Pb, Pd, Pt, Rh, Sb, Se, Si, Sn, Sr and Zn. Furthermore, energy density (Wh/kg) is found by way of the following formula (2). The results are shown in the following Table 2.

$$\text{capacity density} = (\text{specific capacity of positive electrode material} \times \text{specific capacity of negative electrode material}) / (\text{specific capacity of positive electrode material} + \text{specific capacity of negative electrode material}) \dots (1)$$

$$\text{energy density} = \text{capacity density} \times \text{battery voltage} \dots (2)$$

[0061]

(TABLE 2)

	positive electrode material	negative electrode material		capacity density (Ah/kg)	energy density (Wh/kg)
		type	specific capacity (Ah/kg)		
Example A1	$\text{Fe}_2\text{O}_3$	Li-Si	3000	1014	770
Example A2	$\text{Fe}_3\text{O}_4$	Li-Si	3000	757	454
Example A3	$\text{MnFe}_2\text{O}_4$	Li-Si	3000	617	432
Example A4	$\text{NiFe}_2\text{O}_4$	Li-Si	3000	941	564
Example A5	$\text{CoFe}_2\text{O}_4$	Li-Si	3000	945	567
Example A6	$\text{K}_{1.4}\text{Fe}_{11}\text{O}_{17}$	Li-Si	3000	907	544
Comparative Example	$\text{LiCoO}_2$	Graphite	370	107	384

[0062]

As apparent from the results, each of the non-aqueous electrolyte batteries according to Examples A1 to A6 present improved capacity density and energy density than that of Comparative Example.

(Example B1)

[0063]

In Example B1, a positive electrode and a non-aqueous electrolyte that were prepared as follows were used.

(Preparation of Positive Electrode)

[0 0 6 4]

In the preparation of a positive electrode,  $\text{FeS}_2$  was used as a positive electrode material.  $\text{FeS}_2$  as the positive electrode material, and polytetrafluoroethylene as a binding agent were mixed with each other in a ratio of 95 to 5 by weight, were formed by pressure of  $150 \text{ kg/cm}^2$  into a disk of a diameter of 10.3 mm and a weight of 50 mg. The resultant disk was dried in vacuum at  $50^\circ\text{C}$ , thereby to give the positive electrode of Example B1.

(Preparation of Non-aqueous Electrolyte Solution)

[0 0 6 5]

In the preparation of a non-aqueous electrolyte solution, ethylene carbonate and dimethyl carbonate were mixed in a volume ratio of 1 to 1, to give a mixture solvent, in which a solute of hexafluorophosphoric acid lithium ( $\text{LiPF}_6$ ) was dissolved in a concentration of 1 mol/kg.

[0 0 6 6]

Further, in Example B1, as shown in Figure 1, the non-aqueous electrolyte solution 14 prepared as described above was contained in the test cell 10 as well as the positive electrode prepared as described above was used for a working



electrode 11 and lithium metal was used for a counter electrode 12 as a negative electrode and for a reference electrode 13, as in above-mentioned Examples A1 to A6.

[0 0 6 7]

Continuously, the test cell according to Example B1 was subject to discharge at a  $0.5 \text{ mA/cm}^2$  discharging current to a potential of the working electrode 11 to the reference electrode 13 of 1.0 V, followed by charge at a  $0.5 \text{ mA/cm}^2$  charging current to a potential of the working electrode 11 to the reference electrode 13 of 2.5 V.

[0 0 6 8]

With regard to the test cell according to Example B1 wherein  $\text{FeS}_2$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (Ah/kg) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 14.

[0 0 6 9]

As apparent from the results, in the test cell according to Example B1 wherein  $\text{FeS}_2$  is used as the positive electrode material, a flat discharge potential of about 1.5 V is attained, and the specific capacity of the positive electrode material

at the discharge of the first cycle is 524 Ah/kg. Therefore, the positive electrode material presents higher specific capacity, compared with  $\text{LiCoO}_2$  having the specific capacity of about 150 Ah/kg which has been conventionally used as a positive electrode material.

[0070]

Further, in the non-aqueous electrolyte battery according to Example B1 wherein  $\text{FeS}_2$  is used as the positive electrode material, and a lithium alloy (Li-Si) having the specific capacity of 3000 Ah/kg as disclosed in the Japanese Patent Laid-Open No. 2000-321200 and the Japanese Patent Laid-Open No. 2000-321201 is used as the negative electrode material, capacity density (Ah/kg) and energy density (Wh/kg) at the discharge of the first cycle were found in the same manner as in the above-mentioned Examples A1 to A6. The lithium alloy may include Ag, Al, Au, Bi, Ca, Co, Ga, Ge, Hg, In, Ir, Mg, Pb, Pd, Pt, Rh, Sb, Se, Si, Sn, Sr and Zn. The results are shown in the following Table 3.

[0 0 7 1]

(TABLE 3)

	positive electrode material		negative electrode material		capacity density (Ah/kg)	energy density (Wh/kg)
	type	specific capacity (Ah/kg)	type	specific capacity (Ah/kg)		
Example B1	FeS <sub>2</sub>	524	Li-Si	3000	515	670
Comparative Example B1	LiCoO <sub>2</sub>	150	Graphite	370	107	384

[0 0 7 2]

As apparent from the results, the non-aqueous electrolyte battery according to Example B1 presents improved capacity density and energy density, compared with that according to Comparative Example.

(Example C1)

[0 0 7 3]

In Example C1, a positive electrode and a non-aqueous electrolyte solution which were fabricated as followers were used.

(Preparation of Positive Electrode)

[0074]

In the preparation of a positive electrode,  $\text{NaFeO}_2$  which is the transition metal oxide having the crystal structure of space group  $R3m$  is used as a positive electrode material. Further,  $\text{NaFeO}_2$  as the positive electrode material, acetylene black as a conductive agent, and polytetrafluoroethylene as a binding agent were mixed in a ratio of 40 to 40 to 20 by weight, were formed by pressure into a disk of a diameter of 16 mm and a thickness of 0.1 mm. The resultant disk was dried in vacuum at  $110^\circ\text{C}$ , thereby to give the positive electrode of Example C1.

(Preparation of Non-aqueous Electrolyte Solution)

[0075]

In the preparation of a non-aqueous electrolyte solution, ethylene carbonate and dimethyl carbonate were mixed in a volume ratio of 1 to 1, to give a mixture solvent, in which a solute of hexafluorophosphoric acid lithium ( $\text{LiPF}_6$ ) was dissolved in a concentration of 1 mol/kg.

[0076]

Next, in Example C1, as shown in Figure 1, the non-aqueous electrolyte solution 14 prepared as described above was contained in a test cell 10 as well as the positive electrodes prepared as described above was used for a working electrode 11 and lithium metal was used for a counter electrode 12 as

a negative electrode and for a reference electrode 13, as in the above-mentioned Examples A1 to A6.

[0 0 7 7]

Further, the test cell according to Example C1 was subject to discharge at a  $2 \text{ mA/cm}^2$  discharging current to a potential of the working electrode 11 to the reference electrode 13 of 0.5 V, followed by charge at a  $2.0 \text{ mA/cm}^2$  charging current to a potential of the working electrode 11 to the reference electrode 13 of 4.0 V. Afterward, discharge/charge was repeatedly performed as in the above-mentioned manner.

[0 0 7 8]

With regard to the test cell according to Example C1 wherein  $\text{NaFeO}_2$  is used as the positive electrode material, a relationship of a voltage and a specific capacity (mAh/g) of the positive electrode material at a first discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (mAh/g) of the positive electrode material at a first charge is graphed out as a charge curved line described in a solid line in Figure 15. Further, a relationship of a voltage and a specific capacity (mAh/g) of the positive electrode material at a second discharge is graphed out as a discharge curved line described in a broken line while a relationship of a voltage and a specific capacity (mAh/g) of the positive electrode material at a second

charge is graphed out as a charge curved line described in a solid line in Figure 16.

[0 0 7 9]

Discharge/charge was repeatedly performed in a cycle of the above-mentioned discharge/charge process, so that the specific capacity (mAh/g) of the positive electrode material at each cycle was found. The results are shown in the following Figure 17.

[0 0 8 0]

Moreover, a ratio of charge capacity to discharge capacity at each cycle, that is, discharge/charge efficiency (%) was found. The results are shown in the following Figure 18.

[0 0 8 1]

As apparent from the results, in the case where  $\text{NaFeO}_2$ , which is the transition metal oxide having the crystal structure of space group  $R3m$  is used as the positive electrode material, the specific capacity of the positive electrode material at the first discharge is as high as 993 mAh/g. It is also apparent that the specific capacity of the positive electrode material drops sharply between the first and second discharge, and that even after the second discharge at which the specific capacity drops sharply,  $\text{NaFeO}_2$  as the positive electrode material presents extremely higher specific capacity, compared with  $\text{LiCoO}_2$  having the specific capacity of

150 mAh/g which has been conventionally used.

[0082]

In addition, after the second discharge, the positive electrode material presents an extremely small decrease of the specific capacity, and charge efficiency of approximately 100 %, making it possible to charge and discharge stably in high battery capacity.

[0083]

Next, in the non-aqueous electrolyte battery according to Example C1 wherein  $\text{NaFeO}_2$  having the specific capacity of 993 mAh/g is used as the positive electrode material, and a lithium alloy (Li-Si) having the specific material of 3000 mAh/g which is disclosed in the Japanese Patent Laid-Open No. 2000-321200 and the Japanese Patent Laid-Open No. 2000-321201, and in the non-aqueous electrolyte battery according to Comparative Example wherein  $\text{LiCoO}_2$  having the specific capacity of 150 mAh/g is used as the positive electrode material, and graphite having the specific capacity of 370 mAh/g is used as the negative electrode material, capacity density (Ah/kg) and energy density (Wh/kg) at the discharge of the first cycle were found. The results are shown in the following Table 4. The lithium alloy may include Ag, Al, Au, Bi, Ca, Co, Ga, Ge, Hg, In, Ir, Mg, Pb, Pd, Pt, Rh, Sb, Se, Si, Sn, Sr and Zn.

[0084]

(TABLE 4)

	positive electrode material		negative electrode material		capacity density (Ah/kg)	energy density (Wh/kg)
	type	specific capacity (mAh/g)	type	specific capacity (mAh/g)		
Example C1	NaFeO <sub>2</sub>	993	Li-Si	3000	746	448
Comparative Example	LiCoO <sub>2</sub>	150	Graphite	370	107	384

[0085]

As apparent from the results, the non-aqueous electrolyte battery according to Example C1 presents improved capacity density and energy density, compared with that according to Comparative Example.

[0086]

Although the present invention has been fully described by way of examples, it is to be noted that various changes and modification will be apparent to those skilled in the art. Therefore, unless otherwise such changes and modifications depart from the scope of the present invention, they should be constructed as being included therein.